



# SYNTHESIS AND CHARACTERIZATION OF NANOPARTICLES OF 1-HEPTA -O-BENZOYL -B-D-LACTOSYL 5-ARYL 2,4 DITHIOBIURETS

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## Abstract:

In view of application of Nanoparticles and lactosyl compounds in industrial and medicinal research, we here by report the synthesized series of 1-Hepta -O-benzoyl - $\beta$ -D-lactosyl 5-Aryl 2,4 dithiobiurets nanoparticles and are characterized by IR, NMR and X-ray diffractions.

**Key words:** Dithiobiurets, Nanoparticles and characterization.

## Introduction:

Properties of many conventional materials changed when form from nanoparticles. This is typically because nanoparticles have a greater surface area per weight than larger particles which causes them to be more reactive to some other molecule. Carbohydrate especially lactosyl compounds have been uses as starting material in the synthesis of nitrogen and sulphure containing open chain and cyclic compound which was already investigated by earlier workers. Nanoparticles exhibit new physical-chemical properties which are not observed either in individual molecules, or in bulk nanoparticles show unique properties that are significantly different from their bulk materials. In view of this application<sup>1</sup> of lactosyl compounds and Nanoparticles in this we have synthesis to investigate the chemistry of this new compound with reference to their application.

Nanostructure materials are attracting a great deal of attention because of their potential for achieving specific processes and selectivity, especially in biological and pharmaceutical applications<sup>2,3</sup>. Recent studies have demonstrated that especially formulated nanoparticles have good antibacterial activity<sup>4,5</sup>.

## Experimental:

UV-visible Spectra is measured using UV Spectrophotometer by using model Single Beam UV-Visible Spectrophotometer with software(BI/CI/SP/SB-S-03)of Bio Era make. IR spectra were recorded on Perkin-Elmer spectrum RXI FTIR spectrophotometer ( $4000-450\text{ cm}^{-1}$ ).  $^1\text{H}$  NMR was recorded in  $\text{CDCl}_3$  on Bruker DRX-300 spectrometer operating at 300 MHz.

### a) Synthesis of hepta-O-benzoyl- $\alpha$ -D-lactosyl bromide:

The finally powdered lactose octabenzate(0.03M, 21.0g) was added gradually to the brominating agent. After the addition the flask was kept for 2hr at room temperature. Then the reaction mixture with chloroform (130ml) then the mixture was shaken vigorously for about 15 min. The resultant mixture was poured into ice cold water. The chloroform layer was then separated. It was washed several with aqueous sodium bicarbonate to remove excess of acetic acid followed by aqueous sodium metabisulphite to remove excess of bromine and finally 2-3 times with water. To the chloroform addition of petroleum ether afforded a solid (16.5 gm). This solid was expected hepta-O-benzoyl- $\alpha$ -D-lactosyl bromide (yield 77%). It was purified by dissolving it in minimum quantity of chloroform and reprecipitating it with petroleum ether, m.p.  $168^\circ\text{C}$ .

### b) Preparation of lead thiocyanate :

Lead thiocyanate was prepared by mixing aqueous solution of lead nitrate and ammonium thiocyanate. The white granular lead thiocyanate was filtered washed with distilled water and dried at  $50^\circ\text{C}$ .

c) Preparation of hepta-O-benzoyl- $\beta$ -D-lactosyl isothiocyanate<sup>6</sup>:

To a suspension of hepta-O-benzoyl- $\alpha$ -D-lactosyl bromide (21 gm, 0.03M) in sodium dried xylene (80ml) was added lead thiocyanate (6gm, 0.03M). The reaction mixture was then treated for microwave synthesis for about 3 min. This solution was then cooled and liberated lead bromide was removed by filtration. The xylene filtrate was then treated with petroleum ether (60-80°C) with stirring, a white solid mass obtained (13gm). This solid was expected hepta-O-benzoyl- $\beta$ -D-lactosyl isothiocyanate.

It was purified by dissolving it in minimum quantity of chloroform and reprecipitating it with petroleum ether, m.p. 118-120°C. [found; C, 67.07, H, 4.46, N, 1.22, S, 2.9; C<sub>62</sub>H<sub>49</sub>O<sub>17</sub>NS requires; C, 66.96, H, 4.41, N, 1.26, S, 2.88%].

Preparation of 1-hepta O-benzyl- $\beta$ -D-lactosyl 5 phenyl 2,4-Dithiobiurets:

A suspension of 4 gm of Hepta O-benzyl- $\beta$ -D-lactosyl isothiocyanate with 20 ml of benzene and 1 gm of aniline thiourea was treated for microwave synthesis for about 3 min. This IR spectrum of 1-Hepta O-benzyl- $\beta$ -D-lactosyl- 5-phenyl 2,4 Dithiobiurets<sup>7</sup>

solution was then cooled and the benzene filtrate was then treated with petroleum ether (60-80°C) with stirring, a white solid mass obtained (13gm). This solid was expected 1-hepta-O- $\beta$ -D-lactosyl- 5-phenyl 2,4 dithiobiurets.

It was purified by dissolving it in minimum quantity of chloroform and reprecipitating it with petroleum ether, m.p. 145-146°C.

Preparation of Nanoparticles 1-Hepta O-benzyl- $\beta$ -D-lactosyl- 5-phenyl 2,4 Dithiobiurets

Take about 1 gm of 1-Hepta-O-benzyl- $\beta$ -D-lactosyl-5-phenyl 2,4 Dithiobiurets and dissolve complete 1 Hepta O-benzyl- $\beta$ -D-lactosyl- 5-phenyl 2,4 Dithiobiurets in the 50ml of solvent in 250 ml beaker. Now put this beaker in sonicator. The highly penetrating acoustic waves are passed through mixture, which create high pressure bubbles in the beaker due to which breakdown of the bulk material is takes place and desired sized nanoparticles are formed. The size determination of nanoparticles is done by the X-ray diffraction studies.

| Absorption Observed (Cm <sup>-1</sup> ) | Assignment                | Absorption Expected (Cm <sup>-1</sup> ) |
|---|---------------------------|---|
| 3068                                    | C-H Ar-stretching         | 3040-3010                               |
| 1728                                    | C=O stretching            | 1750-1735                               |
| 1176                                    | C-O stretching            | 1210-1153                               |
| 1026, 909                               | Characteristic of lactose | 1100-1000 and 910-900                   |
| 710                                     | Monosubstituted benzene   | 770-680                                 |

NMR SPECTRAL STUDIES<sup>8,9</sup>:

The NMR Spectrum of compound distinctly displayed signals due to N-H Proton at  $\delta$  9.05 and  $\delta$  6.57 ppm, Aromatic Protons at  $\delta$  7.47-7.15 ppm, lactosyl protons at  $\delta$  5.77-3.76 ppm.

Characterisation of Nanoparticles:

1. Characterisation using UV - Visible Spectrophotometer: Characterisation of nanoparticles was done using visible Spectrophotometer by using model Single Beam UV-Visible Spectrophotometer

with software(BI/CI/SP/SB-S-03)of Bio Era make. The UV-Visible Spectroscopy reveals the formation of nanoparticles by showing different absorption those from bulk material.

2. Size determination of Lactose Octabenzoate Nanoparticle by X-Ray Diffraction Studies: From the X-Ray diffraction it comes to know that size of nano octabenzoate is 38nm.

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**References:**

1. [L. Mazur, 2004, *Electrochemistry Communication*, 6, 400-403.]
2. [L. Mazzola, 2003, *Nature Biotechnology*, 21, 1137-1147.]
3. [M. Maillard, S. Giorgio, M. P. Pilent, 2002, *Adv. Material*, 14(15), 1084-1086.]
4. [Z. Chen, L. Gao *Material*, 2007, *Research Bulletin*, 42, 1657-1661.]
5. [C.H.Cao, C.J.Zhou, H.Y.Gao, Y.T. Liu, 2001, *J.Chin.Chem.Soc.*, 48, 207-210.]

6. [P. T. Agrawal, S. P. Deshmukh, 2010, *International Journal of Chem. Tech. Research*, 2(2) 1209 - 1213.]
7. [R.M. Silverstein, G.C. Bassler and B. C. Morrill, "Quantitative Analysis: Identification of Organic Compounds," 2003, 5th Ed., John Wiley and Sons, INC, New York, P 108, 119, 120, 123.]
8. [N.B. Colthup, L.H. Daly and S.E. Wiberley. : *Introduction to Infrared and Raman Spectroscopy*, 2003, Academic Press, New York, P-279.]
9. [D.H. Williams and I. Fleming "Spectroscopic methods in organic chemistry" 4th Ed, Tata McGraw-Hill Publication New Delhi, 2003, 40, 41, 47, 53.]